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The specific heat capacity of a substance is the heat required to increase the temperature of 1g of a substance by 1 o C. The metal can be concluded to have a smaller specific heat than the water because the same amount of energy transfer led to a much larger change in temperature for the metal as compared to the water.

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The Clausius theorem (1855) states that for a thermodynamic system (e.g. heat engine or heat pump) exchanging heat with external reservoirs and undergoing a thermodynamic cycle,  $\oint \frac{\delta Q}{T}$ , where  $\delta Q$  is the infinitesimal amount of heat absorbed by the system from the reservoir and  $T$  is the temperature of the external reservoir (surroundings) at a particular instant in time.